

Group seminar 25, 10, 2019

Effects of Polarity and Polarizability of Solvent on 9,9'-Bifluorenylidene Molecular Balance

Yu Ozawa

Hokkaido University

- located in Sapporo
- 18600 students
- 680 professors
- one Nobel prize



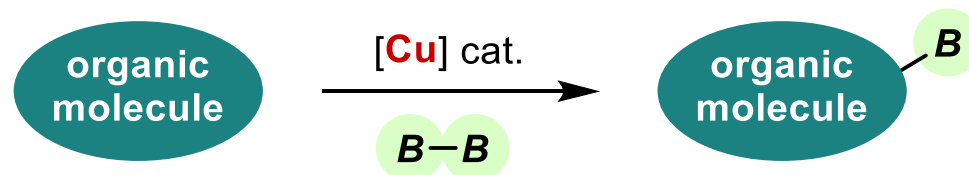
Prof.
Akira Suzuki
(2010)



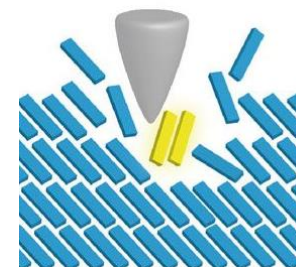
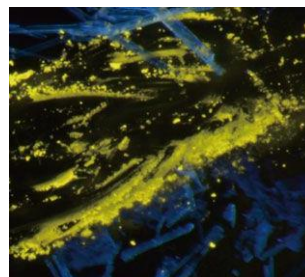
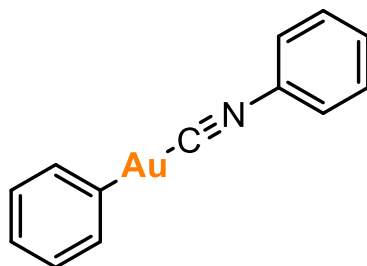


Prof. Hajime Ito

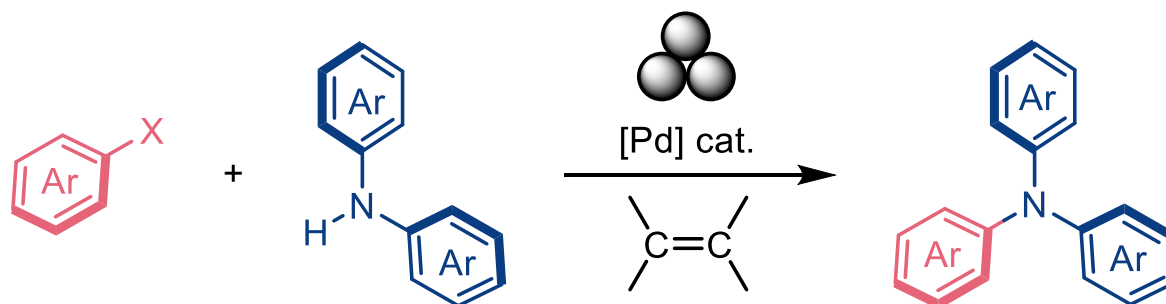
1. Copper(I)-catalyzed borylation reaction

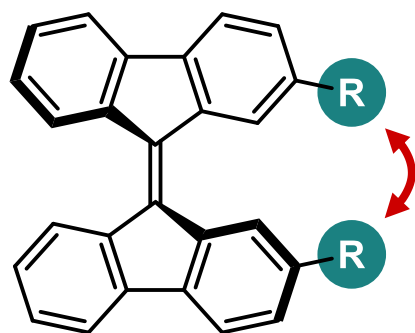


2. Mechano-responsive phosphorescent gold(I) complex



3. Synthetic mechanochemistry

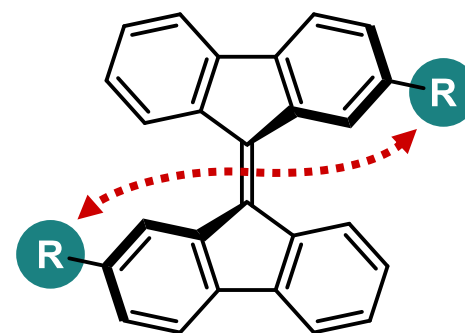
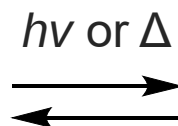




(Z)-1

*Intramolecular
interaction*

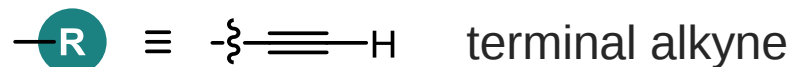
- steric repulsion
- dispersion interaction
- dipole moment
- hydrogen bonding
- C-H/ π interaction
- entropy effect



(E)-1

No interaction

This work



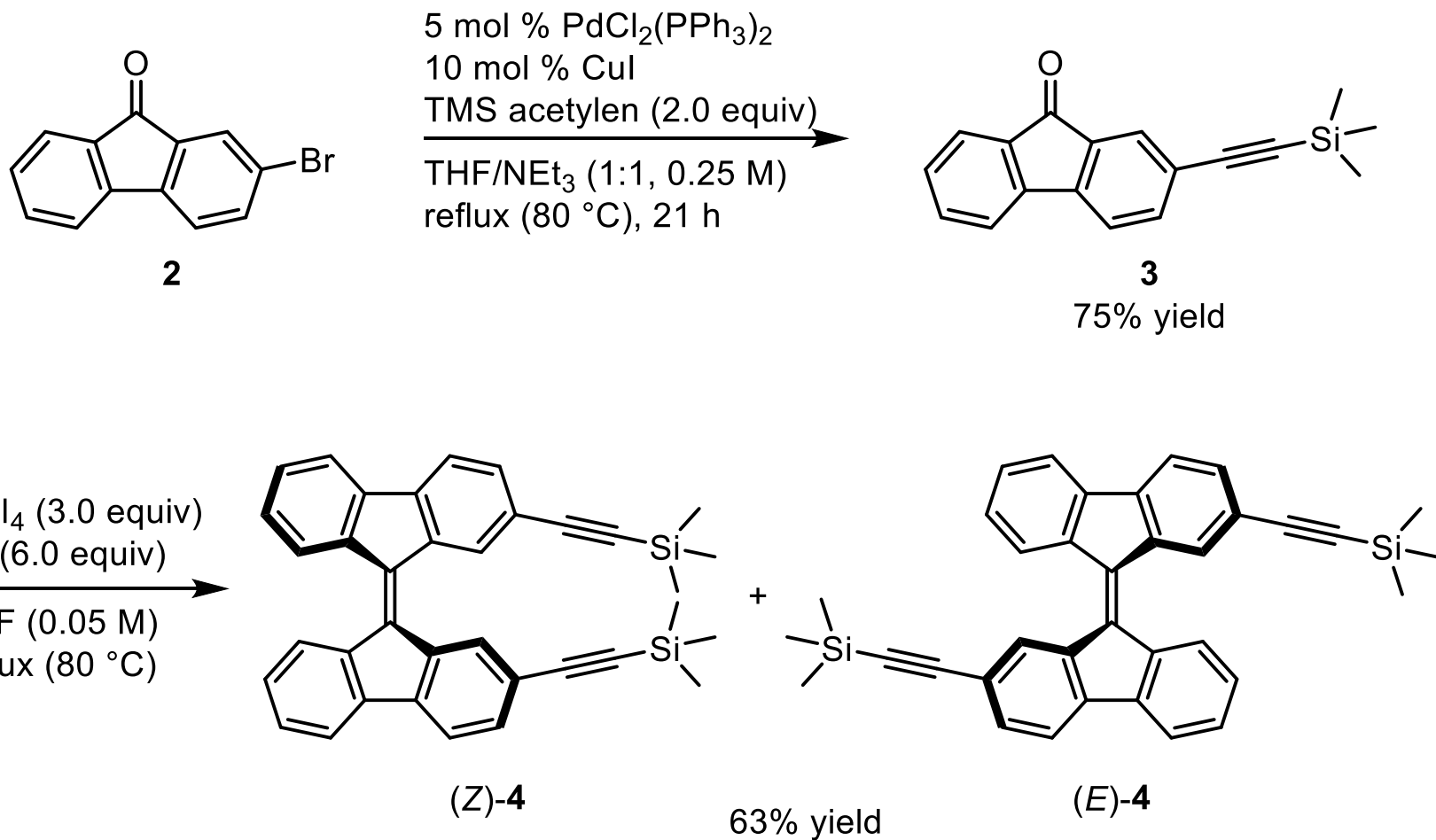
- ▶ smallest functional group in hydrocarbon
- ▶ rigid functional group
- ▶ stand-out dipole moment effect (hypothesis)

Minabe, M.; Oba, T.; Tanaka, M.; Kanno, K.; Tsubota, M. *Chem. Lett.* **2000**, 1, 498.

Minabe, M.; Yamazaki, A.; Imai, T.; Takanezawa, T.; Karikomi, M. *Bull. Chem. Soc. Jpn.* **2006**, 79, 1758.

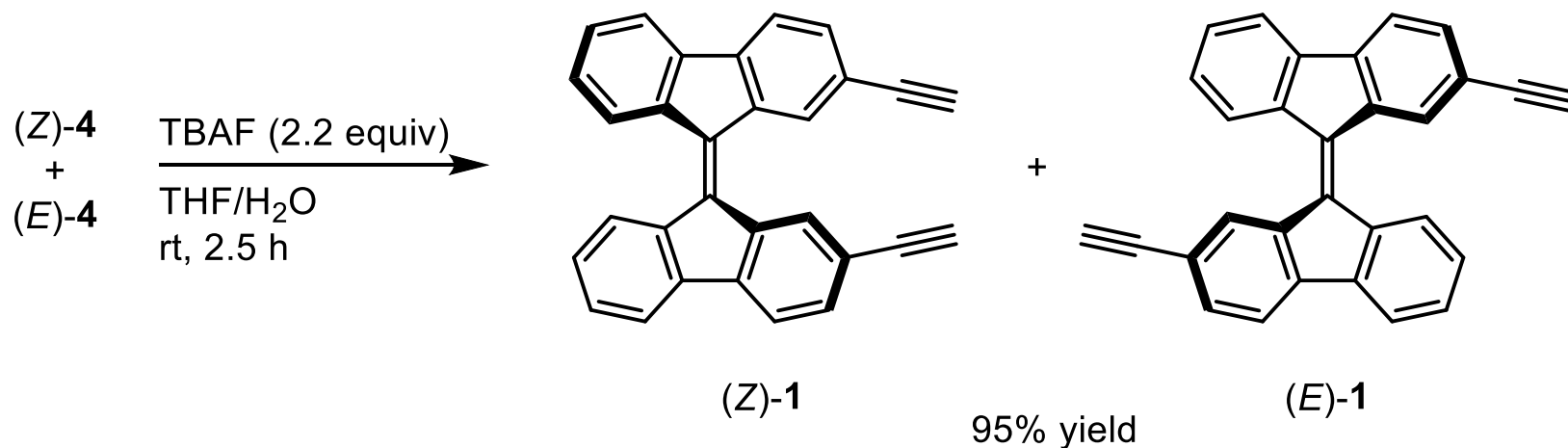
Synthesis of (Z)- and (E)-1

5

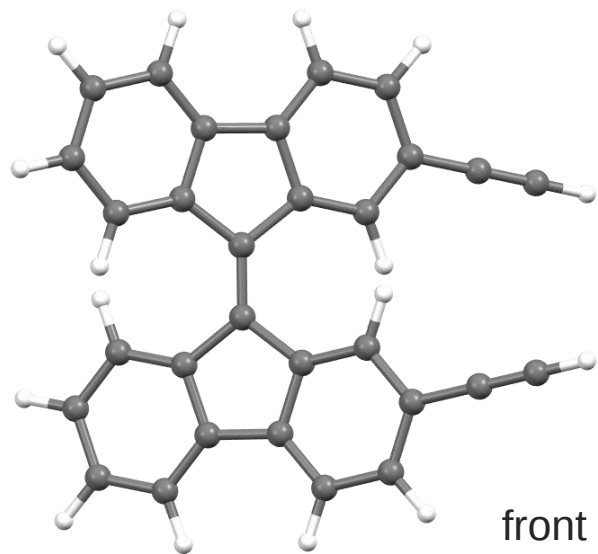


Synthesis of (Z)- and (E)-1

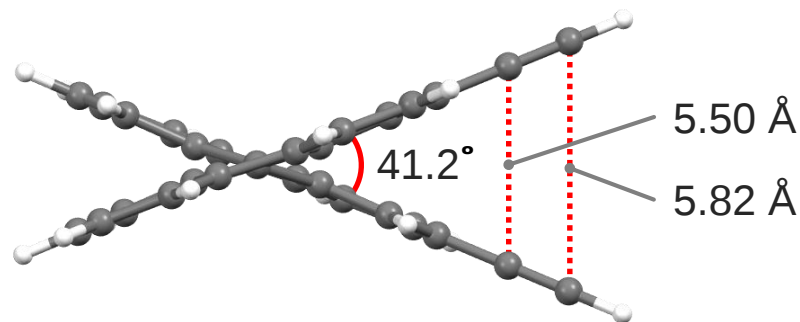
6



Calculated structure of (Z)-1

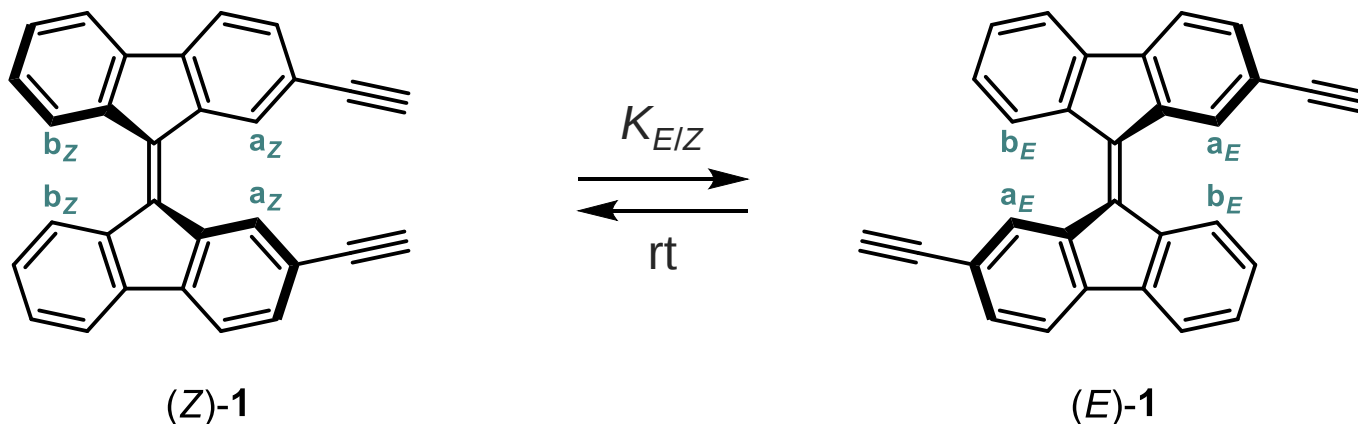


front view



side view

DFT method: B3LYP-D3(BJ)/cc-pVDZ/gas phase.



Definitions

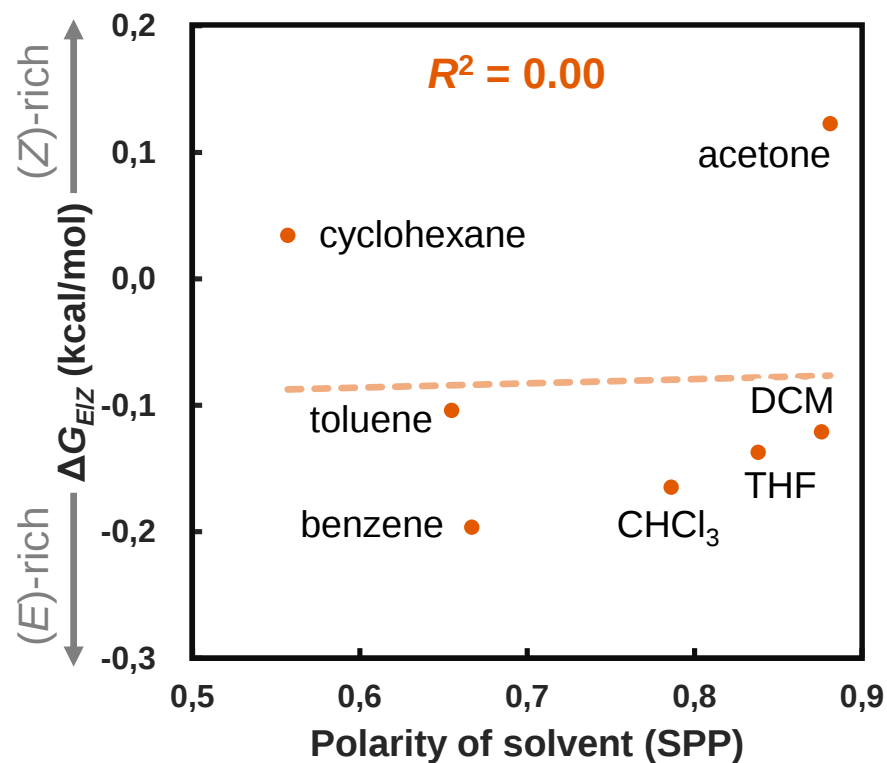
$$K_{E/Z} = \frac{[(E)-1]}{[(Z)-1]}$$

$$\begin{aligned}\Delta G_{E/Z} &= G_E - G_Z \\ &= -RT \ln(K_{E/Z})\end{aligned}$$

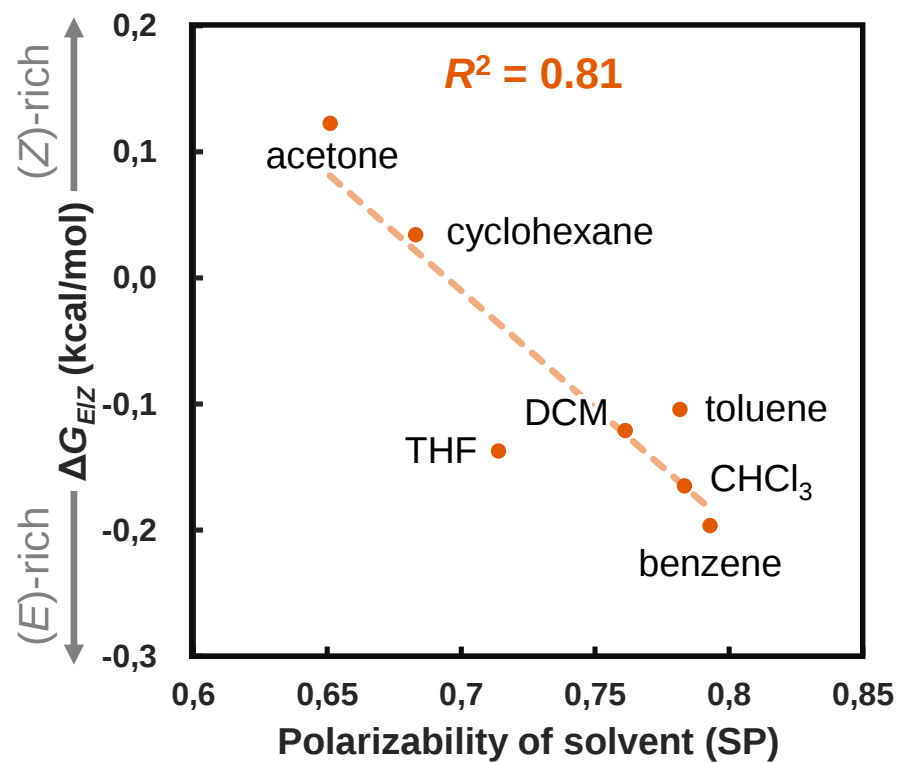
Procedure of NMR study

1. prepare the almost saturated NMR samples.
2. keep the sample dark at r.t. for 24–72 h.
3. take ^1H NMR five times.
4. calculate $K_{E/Z}$ according to the following eq.

$$K_{E/Z} = \frac{\text{Sum of area of } \mathbf{a_E} \text{ and } \mathbf{b_E}}{\text{Sum of area of } \mathbf{a_Z} \text{ and } \mathbf{b_Z}}$$



No correlation

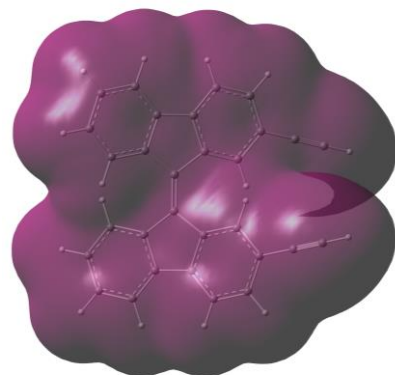


Medium negative correlation

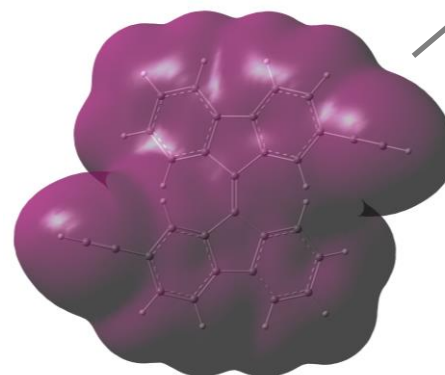
Insight into Effect of Polarizability of Solvent

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1. Difference of solvent contacting surface (ΔS)



(Z)-1
 $\Delta S = 0.00 \text{ \AA}^2$

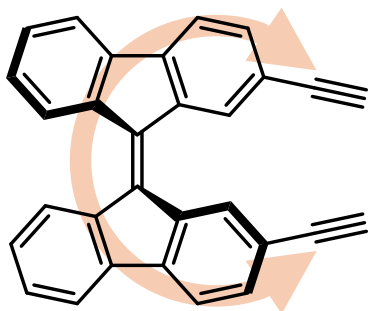


isosurface of
electron density
at 0.0001

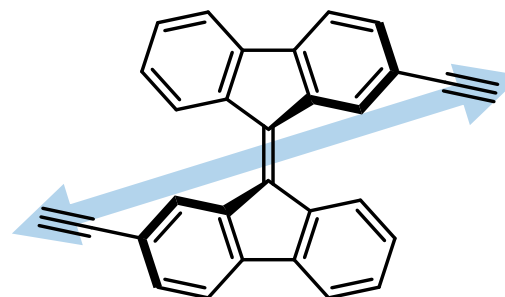
(E)-1
 $\Delta S = 2.55 \text{ \AA}^2$

DFT method: B3LYP-D3(BJ)/Def2-TZVPP/gas phase//B3LYP-D3(BJ)/cc-pVDZ/gas phase.

2. Difference of length of π -system



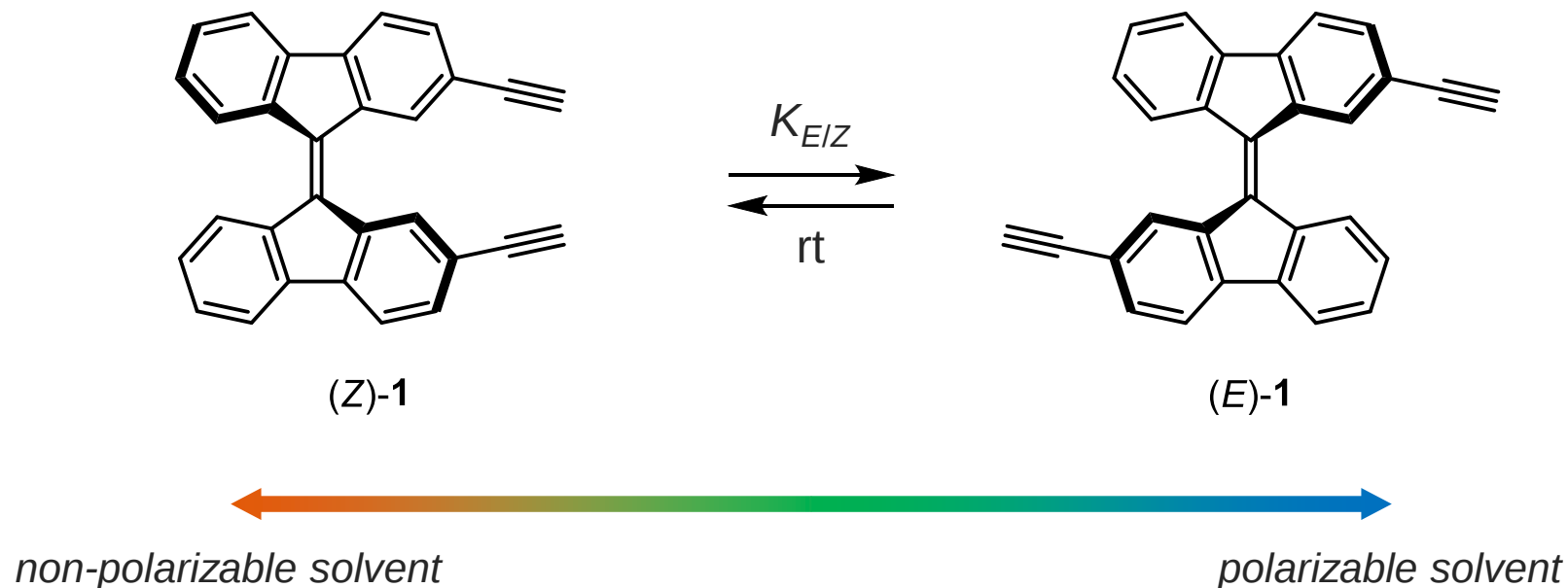
(Z)-1



(E)-1

larger solvent contacting surface
wider π -system

are stabilized in polarizable solvent



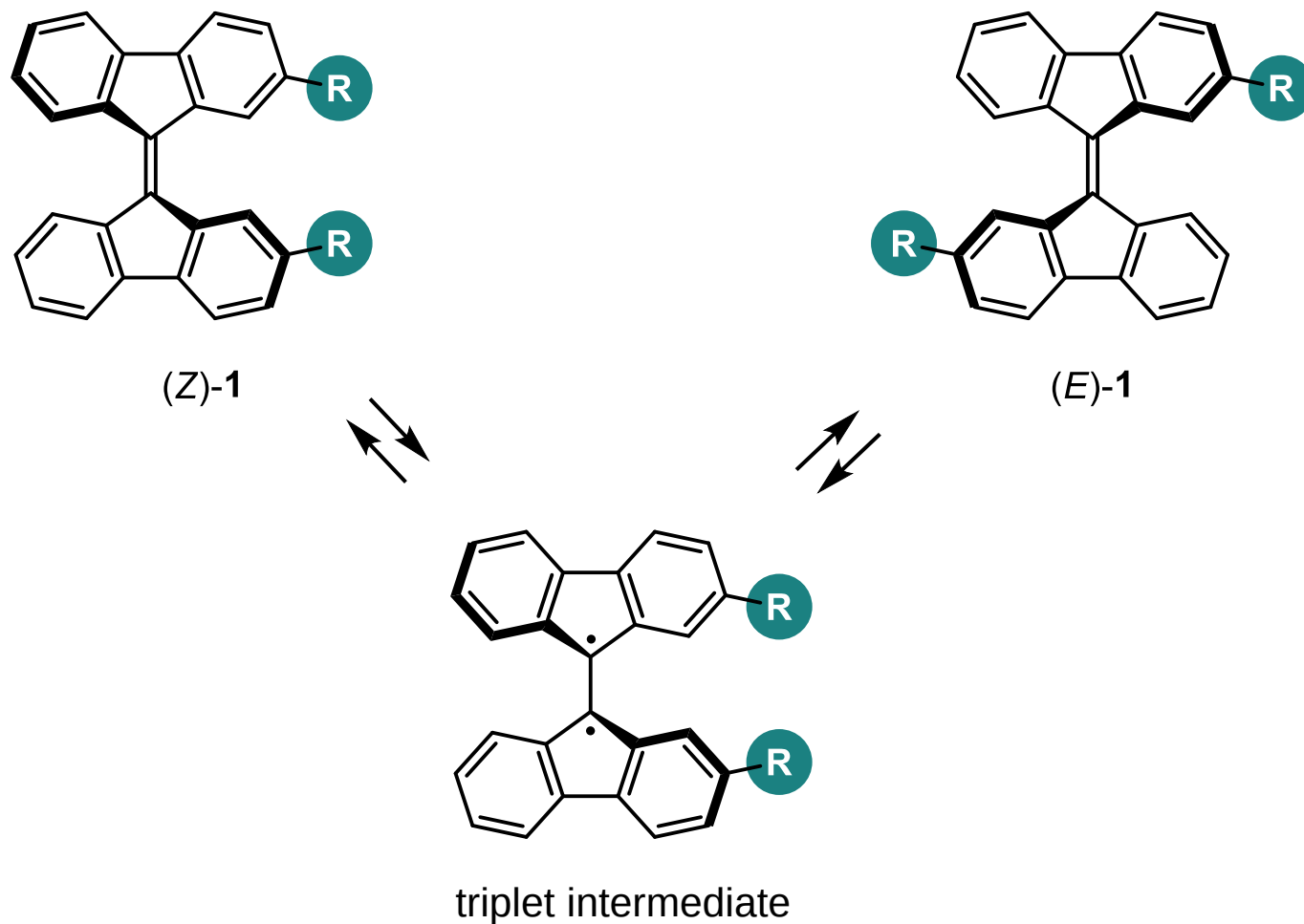
- ▶ polarizability of solvents affects the equilibrium constant $K_{E/Z}$
- ▶ dipole moment of the molecule is not important contrary to expectation
- ▶ dispersion interaction between the molecule and solvents is important

Prof. Peter R. Schreiner

Mr. Finn M. Wilming

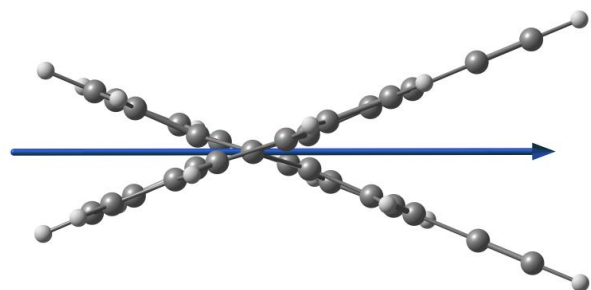
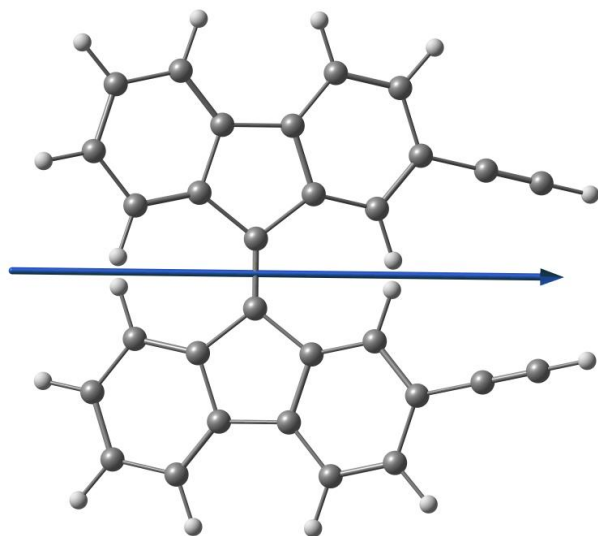
All members in PRS group

Analytics team in Justus Liebig University

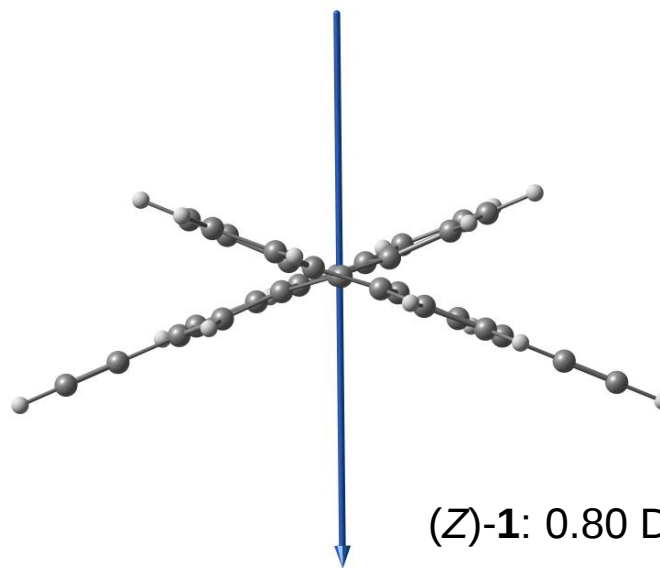
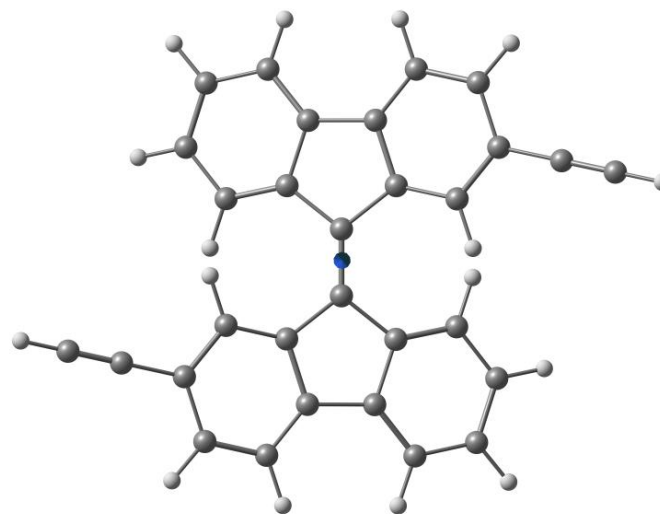


Dipole Moments of (*E*)- and (*Z*)-1

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(*E*)-1: 1.72 Debye

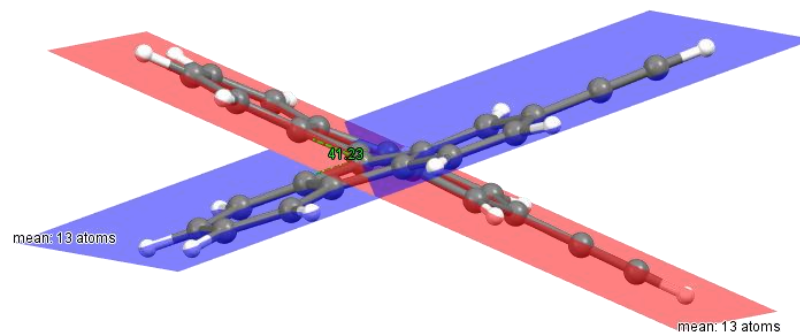
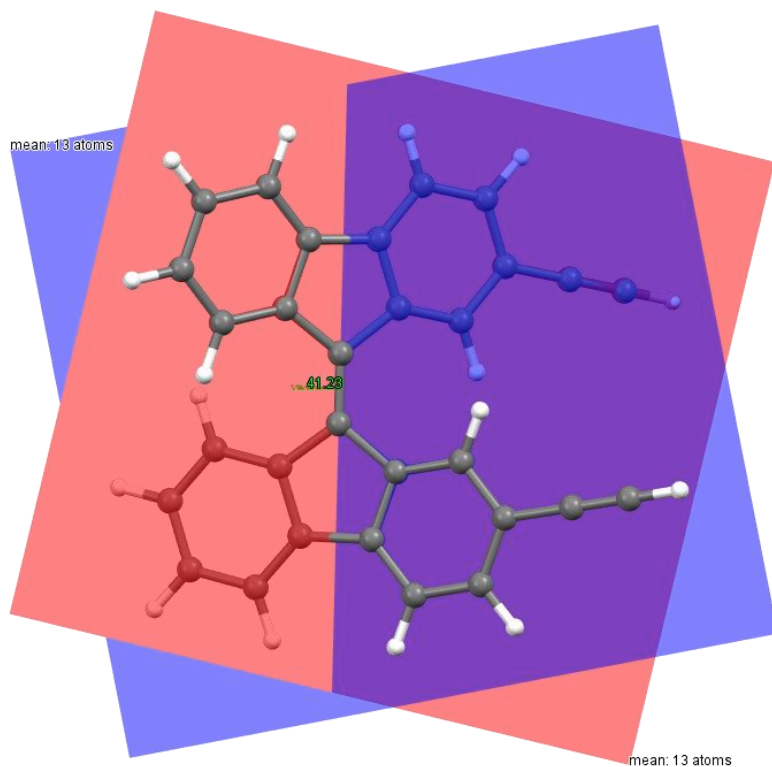


(*Z*)-1: 0.80 Debye

DFT method: B3LYP-D3(BJ)/Def2-TZVPP/gas phase//B3LYP-D3(BJ)/cc-pVDZ/gas phase.

Angle between Two Fluorene Planes in (Z)-1

14

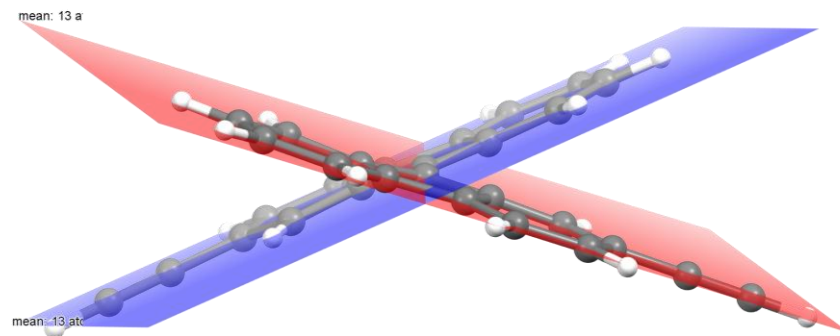
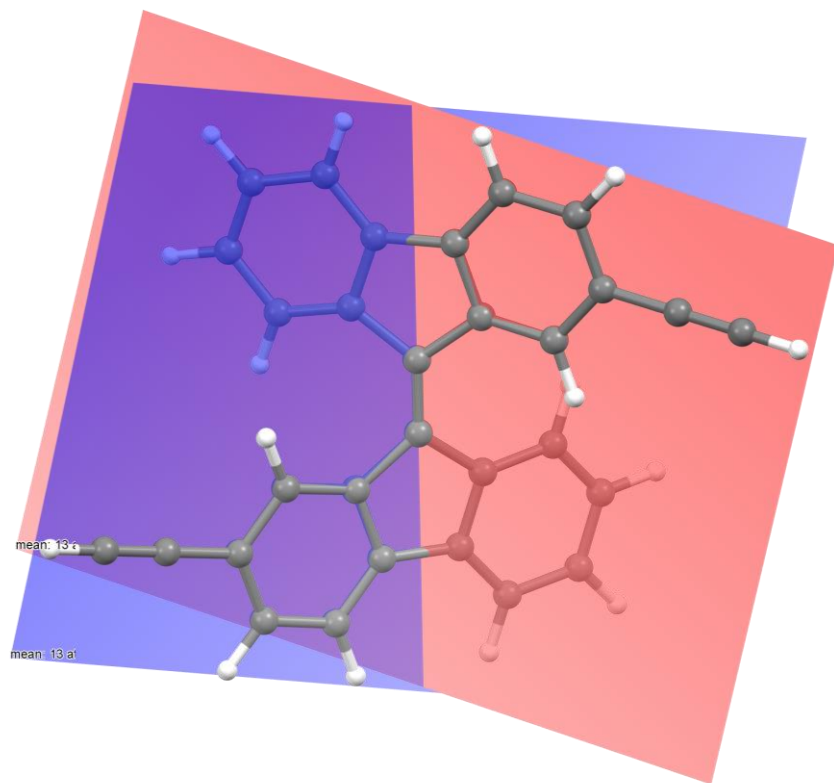


41.2°

DFT method: B3LYP-D3(BJ)/cc-pVDZ/gas phase.

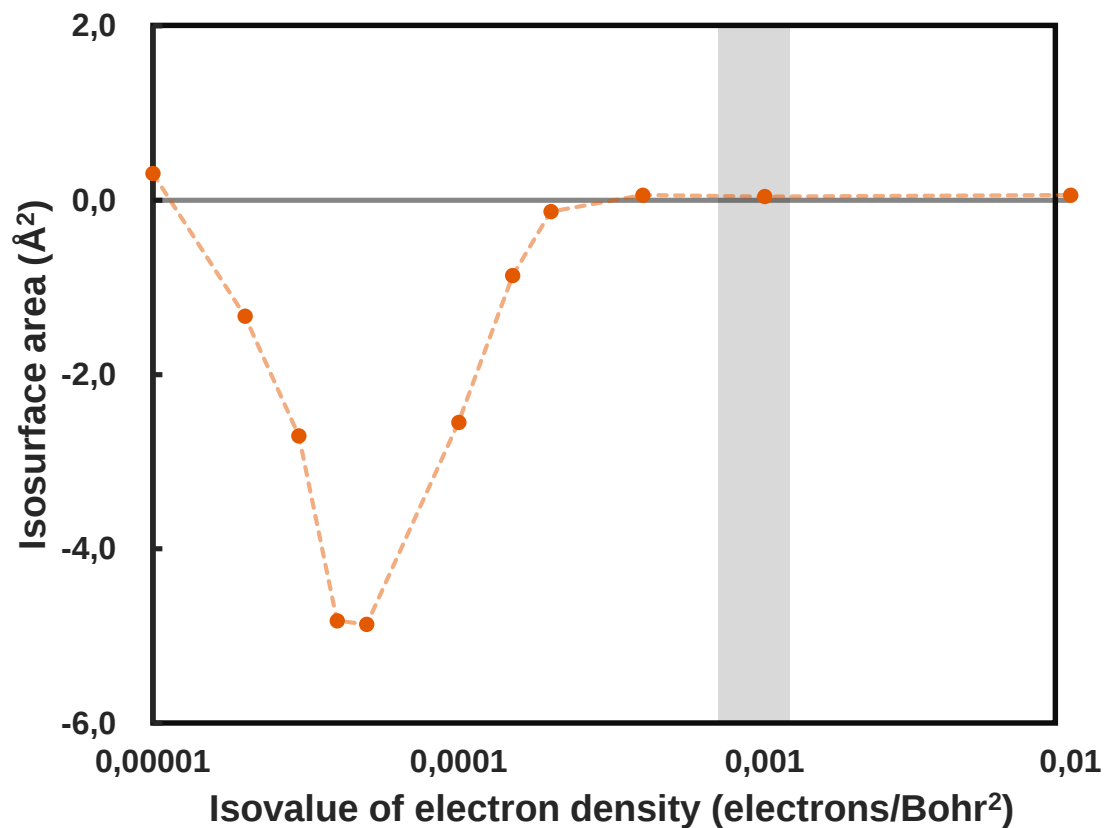
Angle between Two Fluorene Planes in (*E*)-1

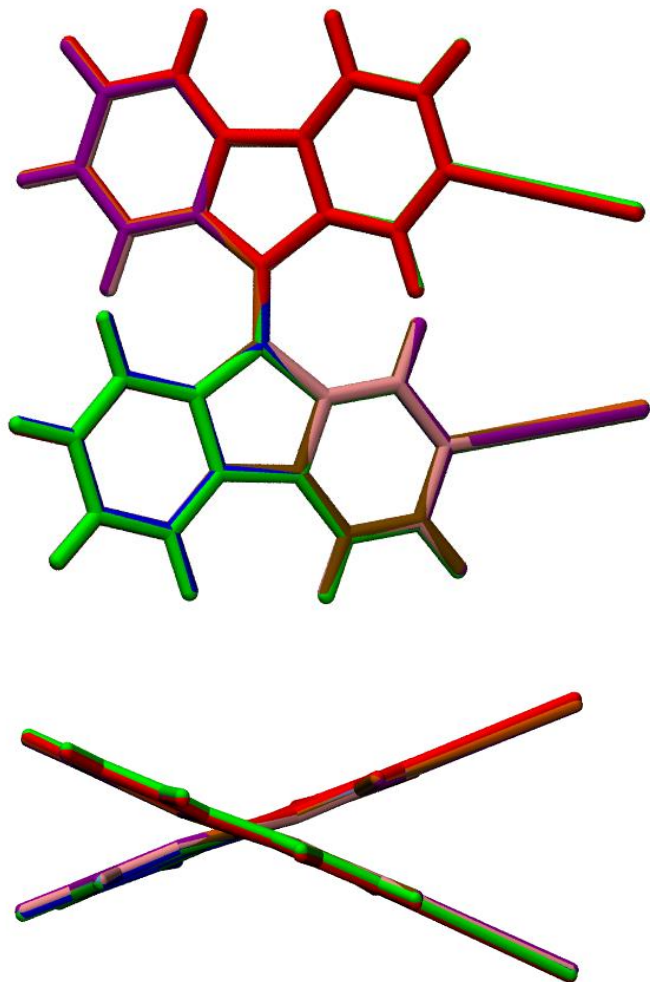
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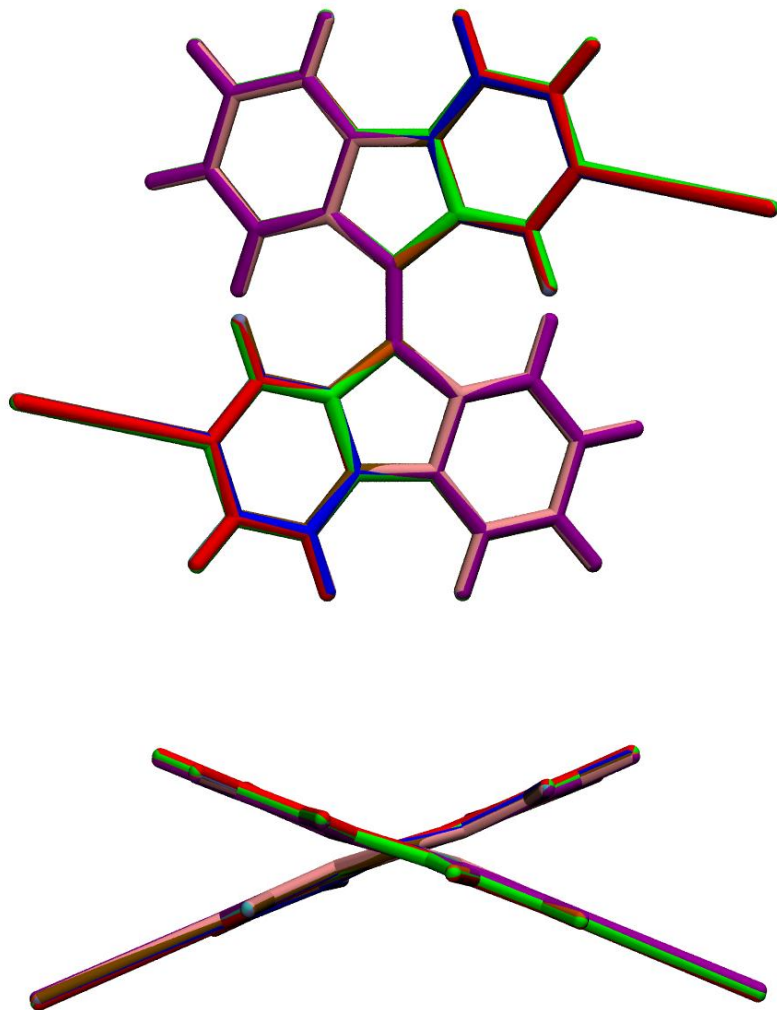
41.1°

DFT method: B3LYP-D3(BJ)/cc-pVDZ/gas phase.





- B3LYP/6-31G(d)
- B3LYP/cc-pVDZ
- B3LYP/Def2-SVP
- B3LYP-D3(BJ)/6-31G(d)
- B3LYP-D3(BJ)/cc-pVDZ
- B3LYP-D3(BJ)/Def2-SVP
- ω B97XD/6-31G(d)
- ω B97XD/cc-pVDZ
- ω B97XD/Def2-SVP
- M06-2X/6-31G(d)
- M06-2X/cc-pVDZ
- M06-2X/Def2-SVP



- B3LYP/6-31G(d)
- B3LYP/cc-pVDZ
- B3LYP/Def2-SVP
- B3LYP-D3(BJ)/6-31G(d)
- B3LYP-D3(BJ)/cc-pVDZ
- B3LYP-D3(BJ)/Def2-SVP
- ω B97XD/6-31G(d)
- ω B97XD/cc-pVDZ
- ω B97XD/Def2-SVP
- M06-2X/6-31G(d)
- M06-2X/cc-pVDZ
- M06-2X/Def2-SVP